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Supplement of

Effects of the 2014 major Baltic inflow on methane and nitrous oxide dynamics in the water column of the central Baltic Sea

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This document contains the following supplementary materials:

Figure S1 – Establishment of the cut-off values below which data for CH₄ and N₂O are indistinguishable from blanks

Text S2 – Precision and accuracy of the reported data

Figure S2 – Diffusion experiment for gas loss from plastic syringes related to text S2.

Figure S3 – Figure exhibiting the changes in water column salinity at the Western Gotland Basin over a period of 16 years.

Table S4 – Table containing detailed information about the specific sampling sites, depths and coordinates.

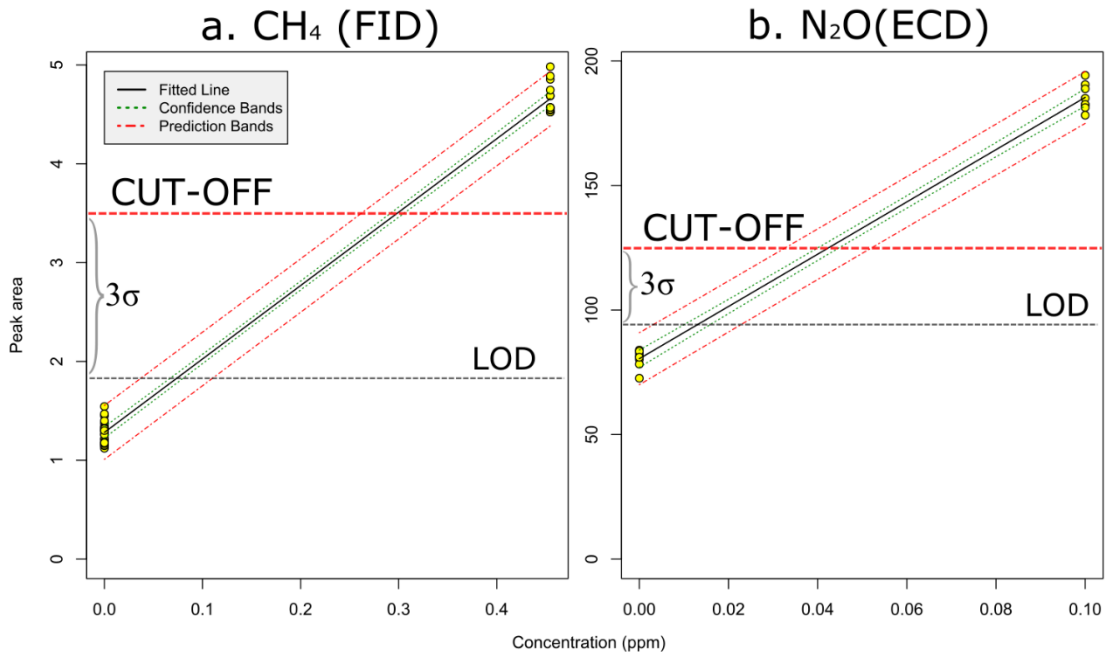


Figure S1. Establishment of cut-off concentrations of CH₄ (left) and N₂O (right), below which measured sample values are regarded as indistinguishable from those of blank pre-evacuated Exetainers. Note the different y-axis scales of the panels, which are due to the different detectors used for the GC measurement of each gas (FID vs. ECD). Yellow points represent measured peak areas of CH₄ and N₂O from blank pre-evacuated Exetainers (i.e. Exetainers from the supplier, subsequently filled with N₂, nominal CH₄ and N₂O concentrations = 0 ppm) and measured peak areas of the lowest standards for each gas, also stored in pre-evacuated Exetainers (CH₄ and N₂O concentrations 0.45 ppm and 0.1 ppm, respectively). A total of 22 replicate blanks and 14 replicate low standards were measured for CH₄ and 8 blanks and 8 low standards for N₂O. The linear fits (peak area vs. ppm) derived from the blanks and the low standards are shown for both gases, with the corresponding 95% confidence and prediction bands. Note that these fits are linear across the full range of standards (not shown). Limit of detection (LOD) was determined according to Armbruster and Pry (2008) ($\text{mean}_{\text{blank}} + 1.645 * \text{SD}_{\text{blank}} + 1.645 * \text{SD}_{\text{low standard}}$). To determine the cut-off, a conservative additional margin was added to the LOD to account for the imprecision of real sample data (in contrast to standard measurements), in which multiple transfers of water and gas between Niskin bottles, syringes and pre-evacuated Exetainers introduce error between replicates. The mean standard deviation in peak area determined from all triplicate samples ($n=N$) was multiplied by 3 and added to LOD to estimate cut-offs of 3.49 and 125.82 peak area units for CH₄ and N₂O, respectively (3σ in the figure). Mean standard deviation of samples was calculated as the square root of mean variances of all sample triplicates: $\sqrt{\frac{\sum \sigma^2_{\text{sample}}}{n_{\text{samples}}}}$.

Text S2. Precision and accuracy of the reported data

Here we estimate the precision and accuracy of the reported data (above the cutoff value), caused by the residual air in the Exetainers, the potential effect of diffusive gas exchange in the plastic syringes and other sources of error, including transfer of samples between syringes.

Precision was estimated from triplicate analysis of all samples and was consistently <5% RSD (see main text). This implies that error introduced by syringe transfers and other sampling artefacts was acceptable for this study where concentrations of CH₄ and N₂O varied over several orders of magnitude.

The average volume of the residual air in the Exetainers was experimentally determined to be 0.7±0.1 mL (n = 20, different batches). The manufacturer (LabCo UK) confirmed that the Exetainers are not flushed with any inert gas prior to evacuation, implying that the residual air has atmospheric concentrations of CH₄ and N₂O (mean concentrations 1.834 ppm, 0.328 ppm, respectively (Blasing, 2016)). This introduces a potential source of inaccuracy in the sample data.

When mixing two varying concentrations of a given gaseous substance at constant temperature, the concentration of the resulting mixture can be calculated by mass balance:

$$C_3 = \frac{(C_1V_1)+(C_2V_2)}{V_1+V_2} \quad (1)$$

The contaminating effect of residual air is expected to be greatest at low sample concentrations, such as those just above the cutoff value (Fig. S1). To estimate the error in accuracy of these samples, let C_1 be the mean atmospheric CH₄ concentration (1.834 ppm), V_1 be the mean volume of residual air in the Exetainers (0.7 mL), V_2 the mean volume of the samples (25 mL) and C_3 the measured mean concentration of CH₄ in the lowest reported samples (approximately 0.4 ppm). By solving eq. 1 for C_2 , which is the true CH₄ concentration in these samples:

$$C_2 = \frac{C_3(V_1+V_2)-C_1V_1}{V_2} \quad (2)$$

it can be calculated that without the contamination, the true mean concentration for a sample measured to contain 0.4 ppm CH₄ was in fact 0.35 ppm. These calculations suggest that the residual air in the Exetainers causes up to an 11.4% overestimation for CH₄ at the low end of our reported range. For N₂O, the overestimation is 10.1% for these samples.

When the sample concentration exceed atmospheric concentration, the effect of the residual air switches from enriching to diluting the true concentrations. Using eq 2. for the highest measured samples in this study (30.1 and 13.2 ppm for CH₄ and N₂O, respectively) there is a 2.6% underestimation in CH₄ and 2.7% for N₂O due to dilution caused by the residual air. The mean concentration in all samples was 2.67±4.11 ppm for CH₄ and 0.24±0.79 ppm for N₂O, meaning that on average, there is a 0.9% underestimation in the CH₄ samples and a 1.1% overestimation in the N₂O samples. We regard these values as acceptable for this study where concentrations of CH₄ and N₂O varied over several orders of magnitude.

The plastic syringes used in sampling (BD Plastipak, 60 mL Luer-Lok) are made of non-gas tight polypropylene, meaning that diffusive exchange with outside air is a potential further source of inaccuracy in the data. To test this, we conducted an experiment in which 15 syringes were filled simultaneously with the same standard gas mixture (5 ppm CH₄, 1.1 ppm N₂O) and subsequently measured after transfer to Exetainers in triplicate, at 30 minute intervals for 150 minutes (Fig. S2). Based on this experiment, we can conclude that the gas loss per hour is < 0.5% for CH₄ and approximately 1.6% for N₂O.

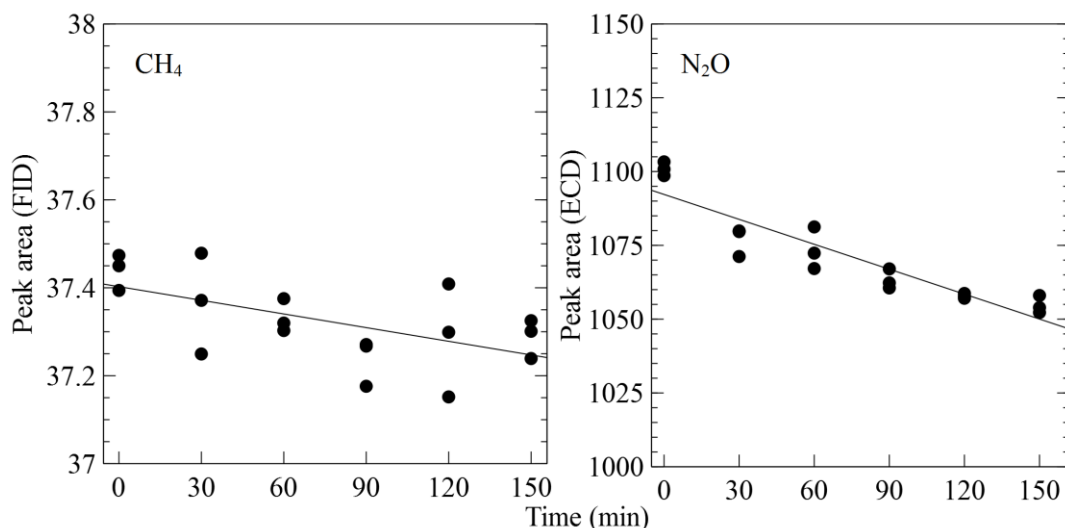


Figure S2. The loss of CH₄ and N₂O from a plastic syringe over 150 minutes in a standard gas mixture containing 5 ppm CH₄ and 1 ppm N₂O. Note the different y-axes.

Thus, given that sample gas remained in the syringes for a maximum of 40 minutes, it can be concluded that diffusive exchange was a relatively small factor in the overall inaccuracy (<1%).

References:

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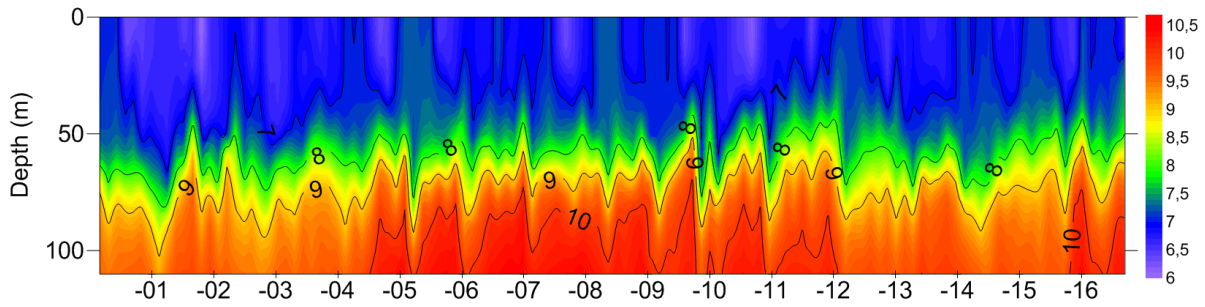


Figure S3. Water column salinity at Western Gotland Basin station BY38 from 2000 to 2016, based on CTD salinity data retrieved from the SMHI Shark database (<http://www.smhi.se/klimatdata/oceanografi/havsmiljodata/marina-miljoovervakningsdata>), exhibiting the typical annual to multi-annual oscillations in halocline depth in this area. These oscillations were likely contributing to the changes in the halocline depth observed during 2015 in the Western Gotland Basin.

